

NAME_____ DATE_____ PERIOD_____
IONIC COMPOUNDS PRACTICE QUIZ #2

CHOOSE THE BEST ANSWER FOR EACH OF THE FOLLOWING

- 1) An ion is an atom or group of atoms which has
a) a positive charge b) a negative charge c) either of these d) no charge
- 2) An ion is an atom or group of atoms which has
a) a different number of protons and electrons
b) the same number of protons and electrons
- 3) A monatomic ion contains
a) only one atom b) 2 or more atoms c) exactly 2 atoms
- 4) A polyatomic ion contains
a) only one atom b) 2 or more atoms c) exactly 2 atoms
- 5) Cations have
a) more electrons than protons b) more protons than electrons c) no charge
- 6) Cations have
a) positive charge b) negative charge c) no charge
- 7) Anions have
a) more electrons than protons b) more protons than electrons c) no charge
- 8) Anions have
a) positive charge b) negative charge c) no charge
- 9) Ionic compounds contain
a) non-metals b) metals c) both of these
- 10) Ionic compounds contain
a) cations b) anions c) both of these
- 11) Covalent compounds contain
a) non-metals b) metals c) both of these

1) COMBINE THE FOLLOWING IONS APPROPRIATELY, AND WRITE THE FORMULA FOR THE RESULTING IONIC COMPOUND

	N^{-3}	Br^{-1}	S^{-2}
Cu^{+2}			
Fe^{+3}			
K^{+1}			
Pb^{+4}			

2) COMBINE THE FOLLOWING IONS APPROPRIATELY, AND WRITE THE FORMULA FOR THE RESULTING IONIC COMPOUND

	PO_4^{-3}	OH^{-1}	CO_3^{-2}
Ca^{+2}			
Al^{+3}			
Na^{+1}			
Sn^{+4}			

WRITE THE FORMULAS FOR THE FOLLOWING **IONIC** COMPOUNDS

3) calcium bromide _____

4) cobalt (II) sulfate _____

5) chromium (III) oxalate _____

6) barium phosphide _____

7) tin (IV) acetate _____

8) iron (II) phosphate _____

9) magnesium cyanide _____

WRITE THE NAMES FOR THE FOLLOWING **IONIC** COMPOUNDS

10) Li_3PO_4 _____

11) Al_2S_3 _____

12) SrCl_2 _____

13) NaCl _____

14) $\text{Ni}_2(\text{SO}_4)_3$ _____

15) $\text{Fe}(\text{MnO}_4)_2$ _____

16) $\text{Cu}_3(\text{PO}_4)_2$ _____

17) K_2SO_3 _____

WRITE THE NAMES FOR THE FOLLOWING **COVALENT** COMPOUNDS

18) NO

19) PF_3

20) N_2O_4

WRITE THE FORMULAS FOR THE FOLLOWING **COVALENT** COMPOUNDS

21) dinitrogen monoxide

22) carbon tetrahydride

23) tetraphosphorus decoxide