

NAME

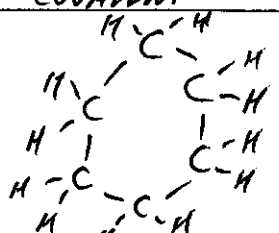
	SO ₂	CO ₂
Ionic or Covalent Compound	COVALENT	COVALENT
Lewis Structure		
Build It (show to your instructor)		
Polar Bonds or Non Polar Bonds	POLAR BOND	POLAR BONDS
Polar Molecule or Non Polar Molecule	POLAR	NON POLAR
Does this compound exhibit dipole forces? -----	YES -----	No -----
If so, does this molecule also exhibit hydrogen bonding?		
Does this compound exhibit LDF's? -----	YES ----- ✓	YES ----- ✓
Which one exhibits greater LDF's?		✓
Which one would evaporate faster?		✓
Which one would have both a higher melting point and higher boiling point?	✓	

[illegible]

	CH ₄	C ₅ H ₁₂
Ionic or Covalent Compound	COVALENT	COVALENT
Lewis Structure	<pre> H H - C - H H </pre>	<pre> H H H H H H - C - C - C - C - C - H H H H H H </pre>
Build It (show to your instructor)		
Polar Bonds or Non Polar Bonds	POLAR BONDS	POLAR BONDS
Polar Molecule or Non Polar Molecule	NON POLAR	NON POLAR
Does this compound exhibit dipole dipole forces?	No	No
If so, does this molecule also exhibit hydrogen bonding?	No	No
Does this compound exhibit LDF's?	YES	YES
Which one exhibits greater LDF's?		✓
Which one would evaporate faster?	✓	
Which one would have both a higher melting point and higher boiling point?		✓

	CH ₃ OH	CH ₂ O
Ionic or Covalent Compound	COVALENT	COVALENT
Lewis Structure	$ \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\bar{\text{O}}-\text{H} \\ \\ \text{H} \end{array} $	$ \begin{array}{c} \text{O} \\ \\ \text{H}-\text{C}-\text{H} \end{array} $
Build It (show to your instructor)		
Polar Bonds or Non Polar Bonds	POLAR BONDS	POLAR BONDS
Polar Molecule or Non Polar Molecule	POLAR	
Does this compound exhibit dipole dipole forces?	YES	YES
If so, does this molecule also exhibit hydrogen bonding?	YES	NO
Does this compound exhibit LDF's?	YES	YES
Which one exhibits greater LDF's?	✓	
Which one would evaporate faster?		✓
Which one would have both a higher melting point and higher boiling point?	✓	

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	CH ₃ OH	C ₆ H ₁₂ (cyclohexane)
Ionic or Covalent Compound	COVALENT	COVALENT
Lewis Structure	$ \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\bar{\text{O}}-\text{H} \\ \\ \text{H} \end{array} $	
Build It (show to your instructor)		
Polar Bonds or Non Polar Bonds	POLAR BONDS	POLAR BONDS
Polar Molecule or Non Polar Molecule	POLAR	NON POLAR
Does this compound exhibit dipole dipole forces?	YES	NO
If so, does this molecule also exhibit hydrogen bonding?	YES	NO
Does this compound exhibit LDF's?	YES	YES
Which one exhibits greater LDF's?		✓
Which one would evaporate faster?	? ✓	? ✓
Which one would have both a higher melting point and higher boiling point?		✓

POLAR ✓
H-BONDS

STRONG
LDF

	Ne	Xe
Ionic or Covalent Compound	COVALENT	COVALENT
Lewis Structure	$\text{I} \ddot{\text{Ne}} \text{I}$	$\text{I} \ddot{\text{Xe}} \text{I}$
Build It (show to your instructor)		
Polar Bonds or Non Polar Bonds	—	—
Polar Molecule or Non Polar Molecule	—	—
Does this compound exhibit dipole dipole forces?	No	No
If so, does this molecule also exhibit hydrogen bonding?	No	No
Does this compound exhibit LDF's?	YES	YES
Which one exhibits greater LDF's?		✓
Which one would evaporate faster?	✓	
Which one would have both a higher melting point and higher boiling point?		✓

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	NaCl	H ₂ O
Ionic or Covalent Compound		
Lewis Structure	$\text{Na}^+ \rightarrow \ddot{\text{Cl}}^-$	$\text{H} - \ddot{\text{O}} - \text{H}$
Build It (show to your instructor)		
Polar Bonds or Non Polar Bonds	IONIC	POLAR BONDS
Polar Molecule or Non Polar Molecule	IONIC	POLAR
Does this compound exhibit dipole dipole forces?	IONIC	YES
If so, does this molecule also exhibit hydrogen bonding?		YES
Does this compound exhibit LDF's?	YES	YES
Which one exhibits greater LDF's?		✓
Which one would evaporate faster?		✓
Which one would have both a higher melting point and higher boiling point?	✓	