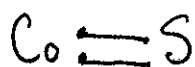


	Chloride Cl^{-1}	Sulfide S^{-2}	Phosphide P^{-3}	Oxide O^{-2}	Iodide I^{-1}	Nitride N^{-3}	Selenide Se^{-2}
Potassium K^{+1}	KCl	K_2S	K_3P	K_2O	KI	K_3N	K_2Se
Cobalt (II) Co^{2+}	CoCl_2	CoS	Co_3P_2	CoO	CoI_2	Co_3N_2	CoSe
Aluminum Al^{3+}	AlCl_3	Al_2S_3	AlP	Al_2O_3	AlI_3	AlN	Al_2Se_3
Sodium Na^{+1}	NaCl	Na_2S	Na_3P	Na_2O	NaI	Na_3N	Na_2Se
Iron (III) Fe^{+3}	FeCl_3	Fe_2S_3	FeP	Fe_2O_3	FeI_3	FeN	Fe_2Se_3
Strontium Sr^{+2}	SrCl_2	SrS	Sr_3P_2	SrO	SrI_2	Sr_3N_2	SrSe
Copper (I) Cu^{+1}	CuCl	Cu_2S	Cu_3P	Cu_2O	CuI	Cu_3N	Cu_2Se_3
Magnesium Mg^{+2}	MgCl_2	MgS	Mg_3P_2	MgO	MgI_2	Mg_3N_2	MgSe
Tin (IV) Sn^{+4}	SnCl_4	SnS_2	Sn_3P_4	SnO_2	SnI_4	Sn_3N_4	SnSe_2
Chromium (III) Cr^{+3}	CrCl_3	Cr_2S_3	CrP	Cr_2O_3	CrI_3	CrN	Cr_2Se_3
Lead (IV) Pb^{+4}	PbCl_4	PbS_2	Pb_3P_4	PbO_2	PbI_4	Pb_3N_4	PbSe_2

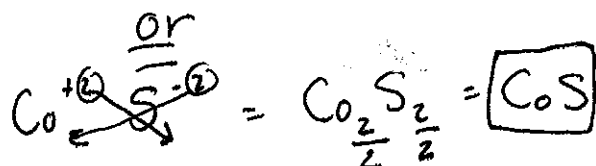
DIRECTIONS: First, write formulas for each ion name in the boxes across the top and down the side. Then, using subscripts, combine the cations and anions using appropriate ratios. Write the formula for each ionic compound in the appropriate box.

$\square$ = Reduced by common factor



2 to lose 2 to gain

You only need 1 Co and 1 S



Same with Sn^{+4} and others w/ +4

$$\text{Sn} \begin{array}{c} \text{---} \\ \text{---} \\ \text{---} \end{array} \text{S}$$

4e- to lose 2e- to gain, need another

So you have SnS_2

