FOSS Mixtures and Solutions Module
Glossary
2005 Edition

Alloy: A combination of two or more metals. (SS)

Atom: The smallest particle of an element. Atoms are the building blocks of matter. (SS)

Catalyst: A chemical that changes the rate of a reaction. (SS)

Change: The process of becoming something different. (TG)

Chemical: A substance used in chemistry. (SS)

Chemical bond: The attachment between two atoms in a molecule. (SS)

Chemical reaction: The process in which two or more substances combine to make one or more new substances that have different properties than the original ones. (TG, SS)

Chemist: A person trained in chemistry. (SS)

Chemistry: The branch of science that deals with the composition, structure, and properties of matter. (SS)

Citric acid: A white, odorless acid with a sour taste. (SS)

Compound: A substance made of two or more elements that are chemically combined. (SS)

Concentration: The relative amount of a substance in a mixture. (SS, TG)

Crystal: The solid form of a material that can be identified by its natural shape or pattern. (TG, SS)

Density: The ratio of the mass of a material in proportion to its volume. (SS)

Dilute: To make a solution less concentrated, usually by adding more liquid. (TG)

Dissolving: The process of a material becoming incorporated uniformly into another or mixing together evenly. (TG, SS)

Element: A substance that cannot be broken down by simple chemical and physical processes. (SS)

Evaporate: To turn into gas, like water into water vapor. (SS)

Evaporation: The process of a liquid turning to gas and dispersing into the air, leaving any dissolved solid material behind. (TG)

Fermenting: A gradual chemical change that takes place without oxygen, caused by organisms such as bacteria or yeast. (SS)

Global warming: Warming of the Earth worldwide. (SS)

Ingot: A mass of metal cast in the shape of a bar. (SS)

Metabolism: The chemical processes that take place in a living cell or organism. (SS)

Mixture: A substance containing two or more materials with different properties. (TG, SS)

Molecule: The smallest part of a substance that is made up of two or more atoms. (SS)

Periodic table: An arrangement of the elements that provides information about their properties. (SS)

Precipitate: A solid material that forms during a chemical reaction. (TG, SS)

Product: A new molecule created in a chemical reaction. (SS)

Property: A characteristic of an object that can be observed, such as size, color, shape, or texture. (TG)

Reactant: A chemical that takes part in a chemical reaction. (SS, TG)

Saturated solution: A solution in which as much solute as possible has been dissolved. (TG, SS)

Solubility: The property substances have of dissolving in solvents, such as the solubility of salt in water. (TG)

Soluble: Capable of being dissolved. Table salt is soluble in water. (SS)

Solute: A substance that dissolves in a solvent to form a solution. (TG, SS)

Solution: A special mixture formed when one or more materials dissolves in another. (TG, SS)

Solvent: A substance that dissolves a solute to form a solution. (TG, SS)

Synthetic: Something made artificially or by a laboratory chemical process, rather than by a natural process. (SS)

Volume: The three-dimensional space occupied by liquid. (TG)

Vulcanization: The process of treating rubber chemically in order to give it useful properties such as strength and temperature stability. (SS)